NAME:_____

Chemistry 150
Stoichiometry
November 3, 2003

1. Over the years, the thermite reaction has been used for welding railroad rails, in incendiary bombs, and to ignite solid fuel rocket motors. The reaction is:

$$Fe_2O_3(s) + 2Al(s) \rightarrow 2Fe(l) + Al_2O_3(s)$$

   a. What masses (in g) of iron(III) oxide and aluminum must be used to produce 15.0 g iron? (5 points)

   b. What is the maximum gram amount of aluminum oxide that could be produced? (5 points)

POINTS:_____(of 10)